

Topic/Objective: States of Matter/predict and describe the motion of particles when thermal energy is added or removed	Name: Ms. H Class/Period: 3 rd & 4 th Block 7 th Grade Date: Tuesday, 10.30.2018	
Essential Question/Phenomena Why does boiling water freeze in extremely, cold dry air?		
Question	Notes	
What type of energy does matter always have?	Key Points Changes in Energy	Information kinetic energy explains Why matter exist
How are the states of matter set apart?	Matter Motion	Solid – low energy and vibrate in one position
What allows the state of matter to be set Apart?		Liquid – have a higher Energy than solid and Move freely over each other
What type of chemical property is Evident in the states of matter? Why?		Gas – very high energy And have no attraction matter is always in constant motion
	Kinetic Theory	particles are attracted to each other but can move Apart with enough energy
How does thermal energy affect the state Of matter from solid to liquid to gas and vice versa?	Thermal Energy	when heat is applied, it speeds particles up Cooling thermal energy decreases the speed Heat is thermal energy transferred from one substance or system to another
Why do solids have a vibration?	Solid	not easily compressed due to the fixed arrangement of particles Particle movement is vibration only, strong bonds
	Liquid	Sublimation – the process of a changing into a gas take the shape of the bottom
SUMMARY: _____ _____ _____ _____ _____		

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Why does liquid take the shape of what Ever container that it is in?	Liquid	of the container More compressible than solids but less compressible than gas Liquid to a gas is called evaporation or vaporization easily compressible Most movement and space of particles More energy (heat) the faster the particles will move Gas to a liquid is condensation, gas to a solid is deposition
	Gas	
	Energy Affects States of Matter	
	Higher Order Thinking Facts	Technilically there's five states of matter: Solid, Liquid, Gas, Plasma, and Bose- Einstein condensates
	Pressure & Matter	Pressure & Matter Diagram
SUMMARY:		
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